FITTING OF POTENTIAL ENERGY SURFACES OF WATER ISOTOPOMERS

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A potential energy surfaces for some major isotopomers of water were constructed by fitting to observed vibration-rotation energy levels of the system using the exact kinetic energy operator nuclear motion program DVR3D [1]. The starting point for the fit is the *ab initi*o Born-Oppenheimer surface and corrections to it: both one- and two-electron relativistic effects, allowance for the Lamb shift and the inclusion of both adiabatic and non-adiabatic non-Born-Oppenheimer corrections [2]. Fits are made by scaling the starting potential by a morphing function, the parameters of which are optimized [3].

Three separate fits were made. Ab initio Born-Oppenheimer surface and corrections to it were optimized using experimental data for $H_2^{16}O$, $H_2^{17}O$ and $H_2^{18}O$ molecules together.

Fit 1 included all data for states with J=0, J=2 and J=5. This constituted 1788 energy levels. From these about 60 were excluded on the grounds that they had a unacceptably large error, in some cases over 1 cm⁻¹. This fit reproduced the data with sigma = 0.071 cm^{-1} .

Fit 2 augmented the data used in Fit 1 with J=10 rotational levels For calculations with J>0 in this fit non-adiabatic effects were used. The best fit was obtained by taking 18% of the values given by Schwenke [4]. Use of the J=10 levels added another 350 levels to the dataset used for Fit 1 and worsened sigma slightly to 0.077 cm⁻¹.

Fit 3 started the data used in Fit 1. From this we excluded all levels which had a residual larger than 0.07 cm^{-1} . A new fit was performed and excluded levels which were accurately reproduced were reintroduced into the fit. This process was repeated with expanded dataset until no further levels could be included without significantly degrading the fit. This fit has a significantly lower sigma of 0.028 cm^{-1} for a dataset of just 1004 levels. This dataset comprises about two-thirds of all the possible levels and spans the entire energy range, up to 25000 cm⁻¹ above the ground state.

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[4]. D.W. Schwenke, J. Chem. Phys., 118, 6898-6904, 2003.